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ADDITIONS & CORRECTIONS

Evidence for Partially Bound States in Cooperative Molecular Recognition Interfaces [*J. Am. Chem. Soc.*, 2008, *130*, 17718−17725] [*J. Am. Chem. Soc.* 2008, *130*, 17718]

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Pages 17722−17723: Equations 3, 4, 6 and 8 relating to the analysis of the sequential binding equilibria are corrected below. The numerical values of the stepwise association constants and effective molarities reported in Tables 5 and 6 are corrected below. Qualitatively, the results are similar to those in the original publication, and the main conclusions are not affected. However, the overall association constant estimated for the doubly H-bonded complex in Table 6 is always larger than the value estimated for a singly H-bonded complex, even when the most populated state of the doubly H-bonded complex is the singly H-bonded species.

$$K_{obs} = K_0(1 + K_1) \tag{3}$$

$$K_{obs} = K_0(1 + K_1 + K_1K_2) \tag{4}$$

$$EM_1 = \frac{K_1}{4K_H} = \frac{K_{obs} - K_0}{4K_0K_H} \tag{6}$$

$$EM_2 = \frac{8K_2}{6K_H} = \frac{8(K_{obs} - K_0 - K_0K_1)}{6K_0K_1K_H} \tag{8}$$

Table 5. Effective Molarities, *EM* (M), Sequential Equilibrium Constants for H-Bond Interactions, K_1 and K_2 (M^{-1}), in the Complexes Formed Between Porphyrin **3** and Ligands **4b** and **4d** at 298 K[a]

solvent	**3·4b**		**3·4d**	
	K_1	EM_1	K_2	EM_2
toluene	14	0.1	14	0.8
TCE	3	0.2	*b*	*b*
DCM	5	0.7	*b*	*b*
$CHCl_3$	1	0.7	*b*	*b*
acetone	5	0.9	*b*	*b*

[a] Errors are EM_1 ± 60%, K_1 ± 40%, EM_2 ± 80% and K_2 ± 60%. [b] The values of K_2 are within experimental error of zero, and so the second H-bond does not confer a measurable additional stability on the complex.

Table 6. Estimated Overall Association Constants for Complexes Formed Between Porphyrin **3** and Ligand **4d** That Make Either One or Two H-bonds (M^{-1}) and Populations of Partially Bound States Assuming $EM_1 = EM_2 = 0.5$ M

solvent	singly H-bonded complex			doubly H-bonded complex			
	K_{est}	zero H-bonds (%)	one H-bond (%)	K_{est}	zero H-bonds (%)	one H-bond (%)	two H-bonds (%)
toluene	4.2×10^4	1	99	8.2×10^5	0	10	90
TCE	2.7×10^3	8	92	1.0×10^4	4	46	50
DCM	5.3×10^3	13	87	1.5×10^4	8	56	35
$CHCl_3$	1.5×10^3	35	65	2.5×10^3	32	58	10
acetone	4.3×10^2	16	84	1.0×10^3	11	59	29

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